



Cambridge International AS & A Level

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CHEMISTRY

9701/44

Paper 4 A Level Structured Questions

May/June 2026

2 hours

You must answer on the question paper.

No additional materials are needed.

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen. Do **not** use correction fluid or tape.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

INFORMATION

- The total mark for this paper is 100.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.
- Important values, constants and standards are printed in the question paper.

This document has **28** pages. Any blank pages are indicated.

- 1 (a) Dilute sodium hydroxide, NaOH(aq) , is electrolysed.

Identify the gas produced at the anode, and write a half-equation for the reaction involved.

gas produced at the anode

half-equation [2]

- (b) Molten lead bromide, $\text{PbBr}_2(\text{l})$, is electrolysed using a current of 2.40 A for 12.0 minutes.

Bromine gas, $\text{Br}_2(\text{g})$, is produced at the anode.

- (i) Calculate the amount, in mol, of $\text{Br}_2(\text{g})$ produced.

amount of $\text{Br}_2(\text{g}) = \dots\dots\dots \text{ mol}$ [2]

- (ii) The $\text{Br}_2(\text{g})$ is collected and its volume is measured at a pressure of 101 kPa and a temperature of 380°C .

Use the ideal gas equation to calculate the volume of $\text{Br}_2(\text{g})$ collected under these conditions.

Give your answer in dm^3 to **three** significant figures.

Show your working.

volume of $\text{Br}_2(\text{g}) = \dots\dots\dots \text{ dm}^3$ [1]



(c) In a different experiment, molten sodium chloride, NaCl(l) , is electrolysed.

36 000 C of charge is transferred. 8.51 g of sodium metal is produced.

Calculate a value for the Avogadro constant, L .

Show your working.

Give your answer to **three** significant figures.

$L = \dots\dots\dots \text{mol}^{-1}$ [2]

[Total: 7]

- 2 (a) Define strong acid and weak acid.

strong acid

weak acid

[2]

- (b) A 1.00 mol dm^{-3} solution of NaOH(aq) is slowly added to a solution containing ethanedioic acid, HOOC-COOH(aq) , until the solution is alkaline.

During the addition of NaOH(aq) , two organic conjugate acid–base pairs are formed in the mixture.

Use the formulas of the species involved to identify the **two** different **organic** conjugate acid–base pairs.

pair 1:
conjugate acid conjugate base

pair 2:
conjugate acid conjugate base

[1]

- (c) A solution of ethanoic acid, $\text{CH}_3\text{COOH(aq)}$, mixed with calcium ethanoate, $\text{Ca(CH}_3\text{COO)}_2\text{(aq)}$, acts as a buffer solution.

- (i) Write an equation for a reaction that has calcium ethanoate as one of the products.

..... [2]

- (ii) Define buffer solution.

.....

.....

..... [1]

- (iii) Write **two** equations to explain how a buffer solution containing $\text{CH}_3\text{COOH(aq)}$ and $\text{Ca(CH}_3\text{COO)}_2\text{(aq)}$ can control pH.

.....

..... [2]



- (d) The pH of a solution of sodium hydroxide, NaOH, is 11.34.

Calculate the pH of a solution of barium hydroxide, Ba(OH)₂, of the same concentration.

pH = [3]

- (e) (i) Define partition coefficient, K_{pc} .

.....
 [1]

- (ii) Equal volumes of butan-1-ol, CH₃CH₂CH₂CH₂OH, and water are added to a separating funnel. 6.00 g of substance **A** is added to this mixture.

The separating funnel is shaken until all of **A** has dissolved. It is left to stand until no further change takes place.

5.32 g of **A** is found to be dissolved in the aqueous layer.

Calculate the partition coefficient, K_{pc} , of **A** between butan-1-ol and water.

Your answer should be less than 1.00.

K_{pc} = [1]

- (iii) A similar experiment is performed to find the partition coefficient, K_{pc} , of **A** between cyclohexane, C₆H₁₂, and water.

Predict whether this K_{pc} is greater than, equal to, or less than your answer to 2(e)(ii).

Explain your answer.

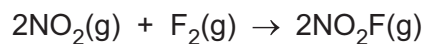
K_{pc} of **A** between C₆H₁₂ and water is my answer to 2(e)(ii).

explanation

 [3]

[Total: 16]

- 3 Nitrogen dioxide reacts with fluorine as shown.



The rate equation for this reaction is shown.

$$\text{rate} = k[\text{NO}_2][\text{F}_2]$$

- (a) Complete Table 3.1.

Table 3.1

order of the reaction with respect to $[\text{NO}_2]$	
order of the reaction with respect to $[\text{F}_2]$	
overall order of the reaction	

[1]

- (b) Table 3.2 shows data from three experiments. The table is incomplete.

Table 3.2

experiment	$[\text{NO}_2]/\text{mol dm}^{-3}$	$[\text{F}_2]/\text{mol dm}^{-3}$	rate / $\text{mol dm}^{-3} \text{s}^{-1}$
1	0.0250	0.0300	3.45×10^{-6}
2	0.0250		1.38×10^{-5}
3	0.0125	0.1200	

- (i) Complete Table 3.2, stating the concentration of F_2 in experiment 2 and the rate of the reaction in experiment 3. [2]
- (ii) Calculate the rate constant, k , under these conditions. Include the units of k .

$k =$

units

[2]



- (c) The experiment is performed again using a large excess of NO_2 .

Under these conditions, the reaction behaves as if it is overall first order. The rate equation is shown.

$$\text{rate} = k'[\text{F}_2]$$

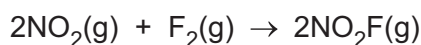
When the initial value of $[\text{F}_2]$ is $4.00 \times 10^{-5} \text{ mol dm}^{-3}$ and the rate of the reaction is measured in $\text{mol dm}^{-3} \text{ s}^{-1}$, the value of k' is 0.0101.

Calculate the half-life of the reaction, $t_{\frac{1}{2}}$, under these conditions.

$$t_{\frac{1}{2}} = \dots \text{ s [1]}$$

- (d) Suggest a two-step mechanism for the reaction between nitrogen dioxide and fluorine.

Identify which of the two steps is the slower step.



$$\text{rate} = k[\text{NO}_2][\text{F}_2]$$

step 1 $\rightarrow$

step 2 $\rightarrow$

slower step [2]

[Total: 8]

4 (a) (i) Define first electron affinity.

.....
..... [1]

(ii) Describe the trend in the first electron affinities of chlorine, bromine and iodine.

Explain your answer.

.....
.....
.....
.....
..... [2]

(b) The lattice enthalpy of potassium bromide, KBr, is -670 kJ mol^{-1} .

Predict values, in kJ mol^{-1} , for the lattice enthalpies of sodium chloride, NaCl, and magnesium oxide, MgO. No calculations are needed.

Explain your predictions.

lattice enthalpy of NaCl = kJ mol^{-1}

explanation
.....
.....

lattice enthalpy of MgO = kJ mol^{-1}

explanation
.....
..... [4]



(c) Magnesium reacts with chlorine as shown.



(i) Define entropy.

.....
 [1]

(ii) Explain why the sign of the entropy change for this reaction is negative.

.....
 [1]

(iii) Predict whether the reaction between magnesium and chlorine becomes more or less spontaneous as temperature increases.

Explain your answer.

The reaction becomes spontaneous.

explanation

 [2]

[Total: 11]



5 (a) The Group 2 nitrates undergo thermal decomposition.

(i) Write an equation for the thermal decomposition of anhydrous calcium nitrate.

..... [1]

(ii) Describe the trend in thermal stability of the nitrate compounds going down Group 2.

Explain your answer.

trend

explanation

.....

.....

.....

.....

[2]



- (b) (i) A spatula measure of magnesium hydroxide, $\text{Mg}(\text{OH})_2(\text{s})$, is stirred into a beaker of $\text{H}_2\text{SO}_4(\text{aq})$. A solution is formed with no solid remaining.

A spatula measure of barium hydroxide, $\text{Ba}(\text{OH})_2(\text{s})$, is stirred into a different beaker of $\text{H}_2\text{SO}_4(\text{aq})$. A permanent precipitate is seen.

Explain these observations.

.....
.....
..... [1]

- (ii) A spatula measure of $\text{Mg}(\text{OH})_2(\text{s})$ is stirred into a beaker of water. Very little of the solid dissolves.

A spatula measure of $\text{Ba}(\text{OH})_2(\text{s})$ is stirred into a different beaker of water. A solution is formed with no solid remaining.

Explain these observations in terms of the energy changes involved.

.....
.....
.....
.....
..... [3]

[Total: 7]

6 Transition elements can form complex ions.

(a) Define complex.

.....
 [1]

(b) Rhodium ions, Rh^{3+} , form ten different octahedral complexes with chloride and ammonia ligands, including isomers.

These complexes have the formula $[\text{Rh}(\text{NH}_3)_x\text{Cl}_y]^n$ where x and y can be any integer from zero to six, and n represents a charge between 3– and 3+.

(i) Complexes **B** and **C** both contain Rh^{3+} , chloride and ammonia ligands.

B has a 2– charge. **C** has a 1+ charge.

State the values of x and y in **B** and in **C**.

x in **B** = y in **B** =

x in **C** = y in **C** =

[2]

(ii) Complex **D**, $[\text{Rh}(\text{NH}_3)_3\text{Cl}_3]$, is uncharged. It exists as a pair of isomers.

State the coordination number of Rh^{3+} in **D**.

..... [1]

(iii) Complete the three-dimensional diagrams in Figure 6.1 to show the two isomers of **D**, $[\text{Rh}(\text{NH}_3)_3\text{Cl}_3]$.

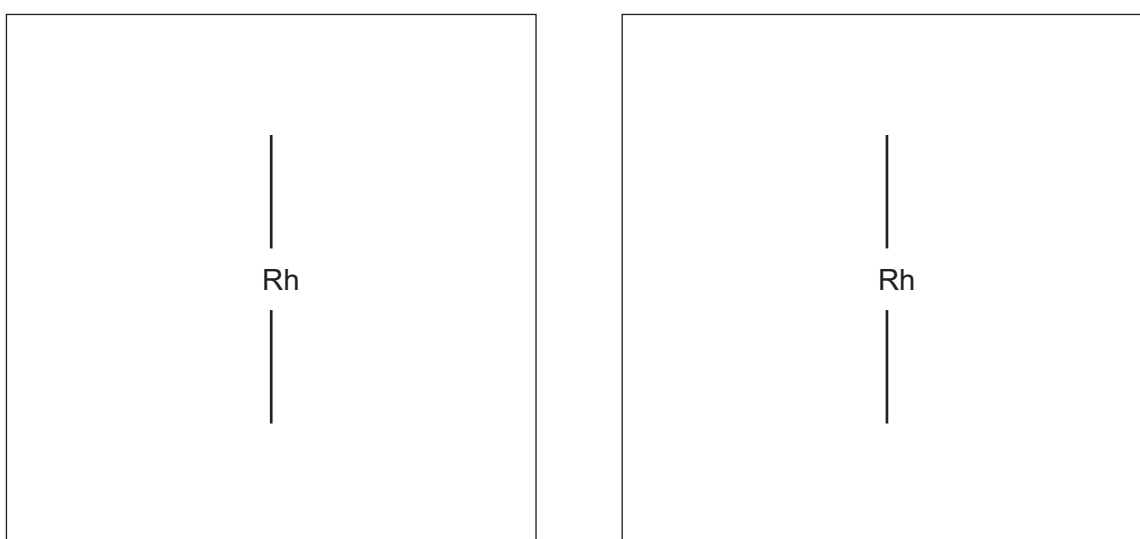


Figure 6.1

[2]



- (iv) Complete Table 6.1 by placing **one** tick (✓) in **each** row to indicate whether the type of isomerism is shown in **D**.

Table 6.1

type of isomerism	shown in D	
	yes	no
stereoisomerism		
optical isomerism		
geometrical isomerism		

[1]

- (c) Nickel(II) salts dissolve in water to form the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex.

Nickel can also form complex **E**, $[\text{Ni}(\text{NH}_3)_6]^{2+}$. The expression for K_{stab} of **E** is shown.

$$K_{\text{stab}} = \frac{[\text{Ni}(\text{NH}_3)_6]^{2+}}{[\text{Ni}(\text{H}_2\text{O})_6]^{2+}[\text{NH}_3]^6} = 4.80 \times 10^7 \text{ (units not shown)}$$

- (i) Define stability constant, K_{stab} .

.....
 [1]

- (ii) State the units of K_{stab} for **E**.

..... [1]

- (iii) **F** is an aqueous solution containing ammonia, NH_3 , **E** and $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$.

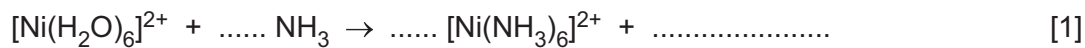
In solution **F**, $[\text{NH}_3] = 0.455 \text{ mol dm}^{-3}$ and $[\text{E}] = 0.0762 \text{ mol dm}^{-3}$.

Calculate the concentration, in mol dm^{-3} , of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ in **F**.

concentration of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ in **F** = mol dm^{-3} [2]



(iv) Complete the equation for the reaction between $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ and NH_3 forming **E**.



(v) Name the type of reaction described in **6(c)(iv)**.

..... [1]

[Total: 13]





Turn over



- 7 (a) The molecular formula of benzene is C_6H_6 . The six carbon atoms in benzene form a six-membered ring.

- (i) State the C–C–H bond angle in benzene and the hybridisation of each carbon atom.

bond angle

hybridisation

[1]

- (ii) The bonds between neighbouring carbon atoms in benzene involve two types of orbital overlap, giving σ and π bonds.

Name the orbitals involved, and describe the orbital overlap, in σ and π bonds.

orbitals involved in σ bonds

orbital overlap in σ bonds

orbitals involved in π bonds

orbital overlap in π bonds

[2]

- (b) Figure 7.1 describes a four-step synthesis of phenol, starting with benzene.

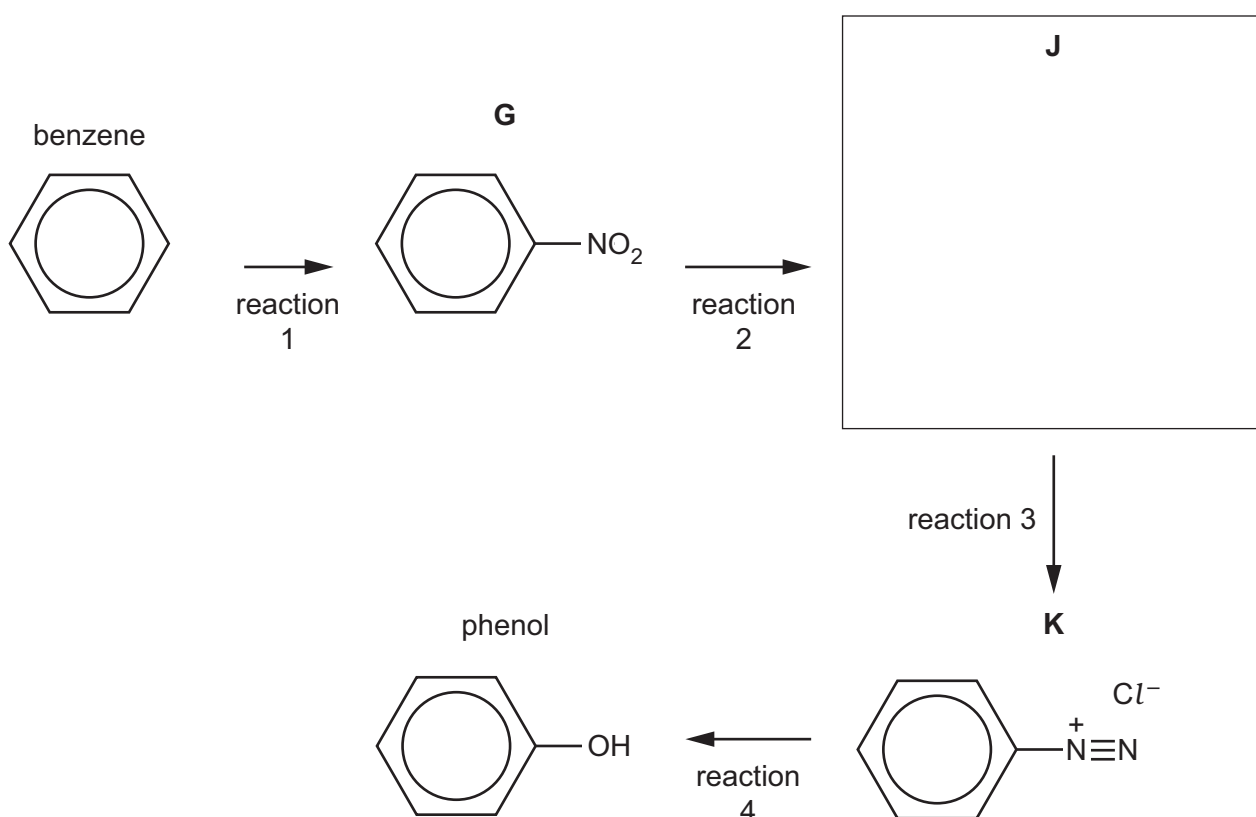


Figure 7.1

- (i) State the reagents and conditions for reaction 1 in Figure 7.1.

..... [1]

- (ii) The reagents and conditions for reaction 2 are concentrated HCl and Sn , heated under reflux.

Draw the structure of compound **J** in Figure 7.1. [1]

- (iii) State the reagents and conditions for reactions 3 and 4 in Figure 7.1.

reaction 3

reaction 4

[2]

- (c) Benzene reacts with ethanoyl chloride, CH_3COCl , in the presence of an AlCl_3 catalyst.

In this reaction, the electrophile is the $\text{CH}_3\text{C}^+\text{O}$ ion.

- (i) Complete the mechanism in Figure 7.2 for this reaction.

Include all relevant curly arrows and charges.

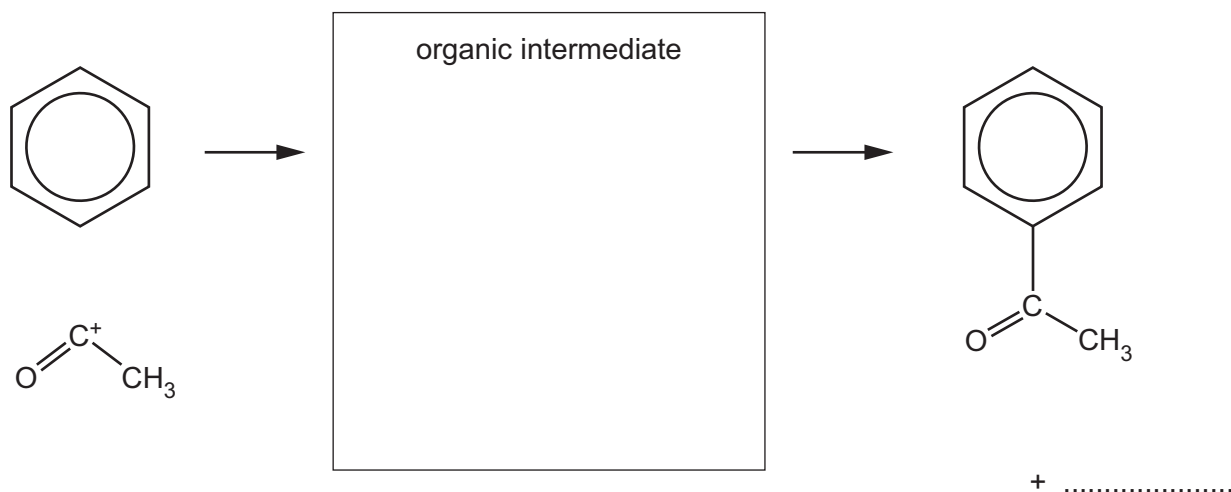


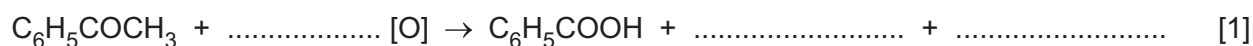
Figure 7.2

[3]

- (ii) The product of the reaction shown in Figure 7.2, $\text{C}_6\text{H}_5\text{COCH}_3$, can be oxidised by heating with alkaline KMnO_4 .

The reaction mixture is then acidified. Benzoic acid, $\text{C}_6\text{H}_5\text{COOH}$, is produced.

Complete the equation for the overall reaction, where $[\text{O}]$ represents one atom of oxygen from the oxidising agent.



(d) Figure 7.3 describes a two-step synthesis of $\text{C}_6\text{H}_5\text{COOH}$.

$\text{C}_6\text{H}_5\text{COCH}_3$ is treated with two reagents, firstly reagent **M** and secondly reagent **N**, producing $\text{C}_6\text{H}_5\text{COOH}$.



Figure 7.3

Identify **M** and **N**, neither contains MnO_4^- or $\text{Cr}_2\text{O}_7^{2-}$.

reagent **M**

reagent **N**

[2]

[Total: 13]





Turn over



- 8 The proton (^1H) NMR spectra of three compounds, **P**, **Q** and **R**, are shown in Figure 8.1. The solvent used is CDCl_3 in each case.

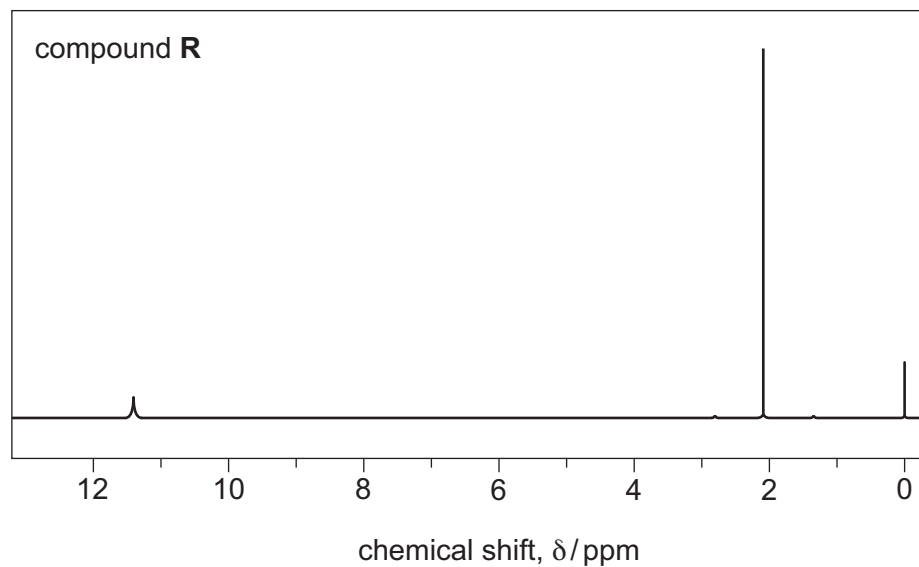
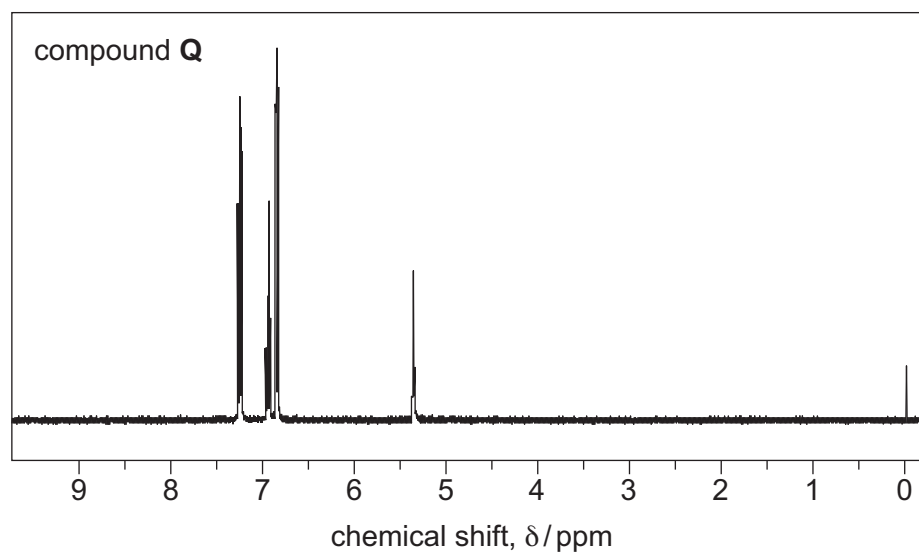
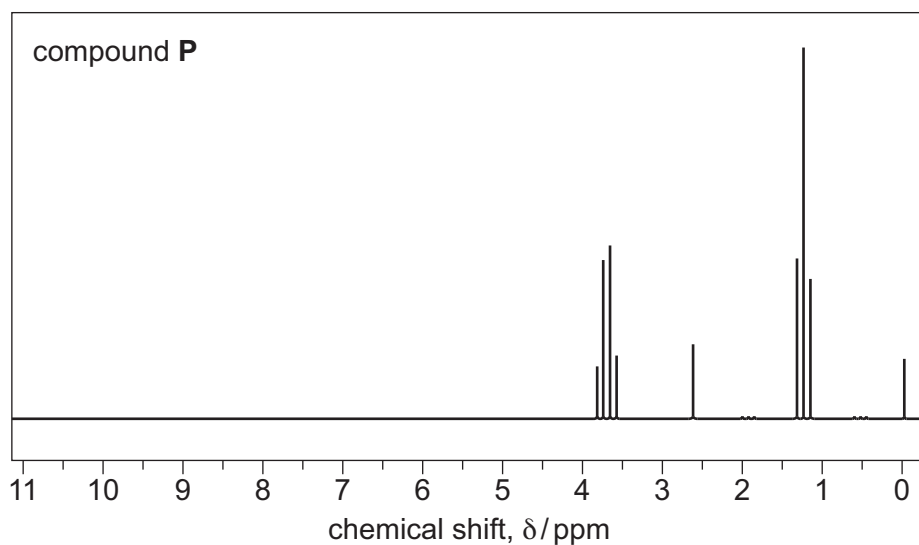


Figure 8.1



Table 8.1

environment of proton	example	chemical shift range δ/ppm
alkane	$-\text{CH}_3$, $-\text{CH}_2-$, $>\text{CH}-$	0.9–1.7
alkyl next to $\text{C}=\text{O}$	$\text{CH}_3-\text{C}=\text{O}$, $-\text{CH}_2-\text{C}=\text{O}$, $>\text{CH}-\text{C}=\text{O}$	2.2–3.0
alkyl next to aromatic ring	CH_3-Ar , $-\text{CH}_2-\text{Ar}$, $>\text{CH}-\text{Ar}$	2.3–3.0
alkyl next to electronegative atom	CH_3-O , $-\text{CH}_2-\text{O}$, $-\text{CH}_2-\text{Cl}$	3.2–4.0
attached to alkene	$=\text{CHR}$	4.5–6.0
attached to aromatic ring	$\text{H}-\text{Ar}$	6.0–9.0
aldehyde	HCOR	9.3–10.5
alcohol	ROH	0.5–6.0
phenol	$\text{Ar}-\text{OH}$	4.5–7.0
carboxylic acid	RCOOH	9.0–13.0
alkyl amine	$\text{R}-\text{NH}-$	1.0–5.0
aryl amine	$\text{Ar}-\text{NH}_2$	3.0–6.0
amide	RCONHR	5.0–12.0

(a) Identify the substance responsible for the peak at $\delta = 0$ ppm on each spectrum.

..... [1]

(b) **P**, **Q** and **R** are phenol, ethanol and ethanoic acid, but not necessarily in this order.

(i) Identify which compound is phenol.

Explain your answer, identifying the protons responsible for the peaks in the relevant spectrum.

phenol is compound

explanation

.....

.....

[2]



- (ii) The spectrum of **P** has peaks at $\delta = 1.2$ ppm and $\delta = 3.7$ ppm.

Name the splitting pattern of each of these peaks and give a reason why the peaks are split in this way.

splitting pattern of the peak at $\delta = 1.2$ ppm

reason for this splitting pattern

.....

splitting pattern of the peak at $\delta = 3.7$ ppm

reason for this splitting pattern

.....

[2]

- (c) The spectrum of **Q** dissolved in D_2O is obtained.

Describe how this spectrum differs from the spectrum of **Q** shown in Figure 8.1.

Explain your answer.

difference

explanation

[1]

- (d) Compare the acidities of phenol, ethanol and ethanoic acid.

Explain your answer.

.....

most acidic

least acidic

explanation

.....

.....

.....

.....

.....

[4]

[Total: 10]



- 9 (a) Using a gel buffered at pH 7.6, electrophoresis is performed on a mixture of three amino acids: asparagine, histidine and lysine.

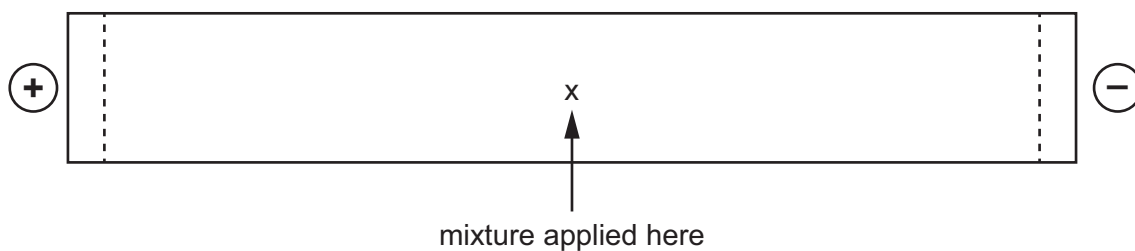


Figure 9.1

The isoelectric points of these amino acids are given in Table 9.1.

Table 9.1

amino acid	isoelectric point
asparagine	5.4
histidine	7.6
lysine	9.7

Describe the direction of movement of each amino acid during the electrophoresis.

Explain your answers.

asparagine

.....

histidine

.....

lysine

.....

[3]





(b) Serine, $\text{H}_2\text{NCH}(\text{CH}_2\text{OH})\text{COOH}$, is an amino acid. It exists as a pair of optical isomers.

- (i) Draw diagrams in Figure 9.2 to show the three-dimensional structures of the two optical isomers of serine.

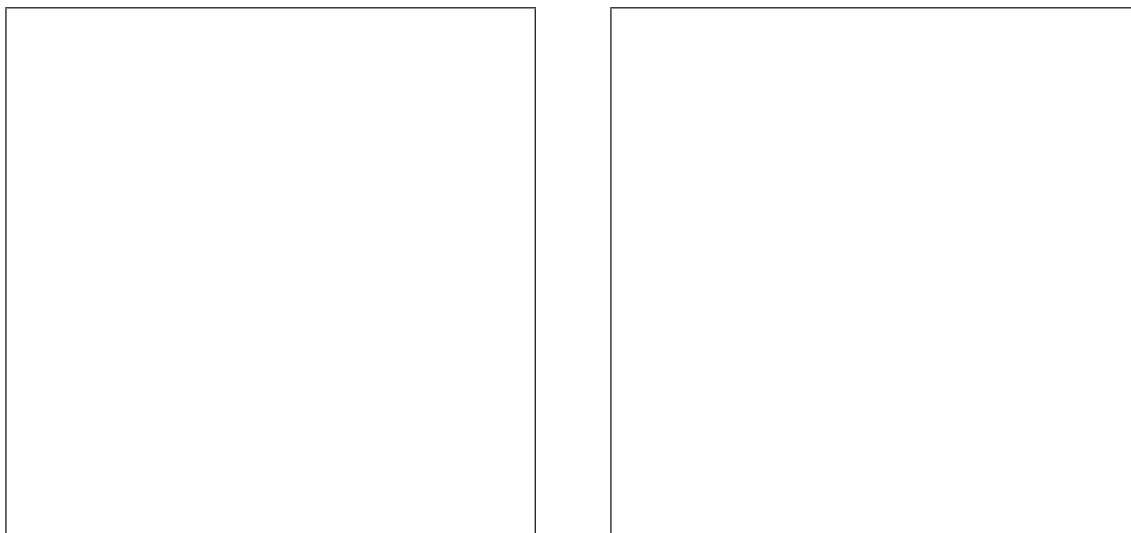


Figure 9.2

[1]

- (ii) Give the name for a mixture containing equal amounts of the two optical isomers.

..... [1]

- (iii) Describe **two** properties that are different for the two optical isomers.

.....

.....

..... [2]

- (iv) Serine is treated with an excess of phosphorus pentachloride, PCl_5 .

The phosphorus-containing product of this reaction is POCl_3 .

Complete the equation for this reaction. Use the structural formula of the organic product.



..... [2]

[Total: 9]



10 ClOCCOCl can be used to make many polymers.

(a) Suggest the name of ClOCCOCl .

..... [1]

(b) (i) ClOCCOCl reacts with $\text{H}_2\text{N}(\text{CH}_2)_4\text{NH}_2$ to form a polymer.

Draw a section of the polymer showing only **one** repeat unit.

Any functional groups should be fully displayed.

[2]

(ii) State the type of polymerisation involved and the linkage between the monomers.

type of polymerisation

linkage

[1]

(c) When $\text{H}_2\text{NCO}(\text{CH}_2)_2\text{CONH}_2$ reacts with a suitable reagent, $\text{H}_2\text{N}(\text{CH}_2)_4\text{NH}_2$ forms.

Identify this reagent.

..... [1]

(d) 1,4-dibromobutane can be used to form $\text{H}_2\text{N}(\text{CH}_2)_4\text{NH}_2$.

Describe the reagents and conditions used.

.....

..... [1]

[Total: 6]





Important values, constants and standards

molar gas constant	$R = 8.31 \text{ J K}^{-1} \text{ mol}^{-1}$
Faraday constant	$F = 9.65 \times 10^4 \text{ C mol}^{-1}$
Avogadro constant	$L = 6.022 \times 10^{23} \text{ mol}^{-1}$
electronic charge	$e = -1.60 \times 10^{-19} \text{ C}$
molar volume of gas	$V_m = 22.4 \text{ dm}^3 \text{ mol}^{-1}$ at s.t.p. (101 kPa and 273 K) $V_m = 24.0 \text{ dm}^3 \text{ mol}^{-1}$ at room conditions
ionic product of water	$K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ (at 298 K (25 °C))
specific heat capacity of water	$c = 4.18 \text{ kJ kg}^{-1} \text{ K}^{-1}$ (4.18 J g ⁻¹ K ⁻¹)

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The Periodic Table of Elements

Group																							
1	2	1 H hydrogen 1.0														13	14	15	16	17	18		
3 Li lithium 6.9	4 Be beryllium 9.0	Key atomic number atomic symbol name relative atomic mass																5 B boron 10.8	6 C carbon 12.0	7 N nitrogen 14.0	8 O oxygen 16.0	9 F fluorine 19.0	10 Ne neon 20.2
11 Na sodium 23.0	12 Mg magnesium 24.3																	13 Al aluminium 27.0	14 Si silicon 28.1	15 P phosphorus 31.0	16 S sulfur 32.1	17 Cl chlorine 35.5	18 Ar argon 39.9
19 K potassium 39.1	20 Ca calcium 40.1	21 Sc scandium 45.0	22 Ti titanium 47.9	23 V vanadium 50.9	24 Cr chromium 52.0	25 Mn manganese 54.9	26 Fe iron 55.8	27 Co cobalt 58.9	28 Ni nickel 58.7	29 Cu copper 63.5	30 Zn zinc 65.4	31 Ga gallium 69.7	32 Ge germanium 72.6	33 As arsenic 74.9	34 Se selenium 79.0	35 Br bromine 79.9	36 Kr krypton 83.8						
37 Rb rubidium 85.5	38 Sr strontium 87.6	39 Y yttrium 88.9	40 Zr zirconium 91.2	41 Nb niobium 92.9	42 Mo molybdenum 95.9	43 Tc technetium —	44 Ru ruthenium 101.1	45 Rh rhodium 102.9	46 Pd palladium 106.4	47 Ag silver 107.9	48 Cd cadmium 112.4	49 In indium 114.8	50 Sn tin 118.7	51 Sb antimony 121.8	52 Te tellurium 127.6	53 I iodine 126.9	54 Xe xenon 131.3						
55 Cs caesium 132.9	56 Ba barium 137.3	57–71 lanthanoids		72 Hf hafnium 178.5	73 Ta tantalum 180.9	74 W tungsten 183.8	75 Re rhenium 186.2	76 Os osmium 190.2	77 Ir iridium 192.2	78 Pt platinum 195.1	79 Au gold 197.0	80 Hg mercury 200.6	81 Tl thallium 204.4	82 Pb lead 207.2	83 Bi bismuth 209.0	84 Po polonium —	85 At astatine —	86 Rn radon —					
87 Fr francium —	88 Ra radium —	89–103 actinoids		104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	113 Nh nihonium —	114 Fl flerovium —	115 Mc moscovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganeson —					

lanthanoids

57	La lanthanum 138.9	58	Ce cerium 140.1	59	Pr praseodymium 140.9	60	Nd neodymium 144.2	61	Pm promethium	62	Sm samarium 150.4	63	Eu europium 152.0	64	Gd gadolinium 157.3	65	Tb terbium 158.9	66	Dy dysprosium 162.5	67	Ho holmium 164.9	68	Er erbium 167.3	69	Tm thulium 168.9	70	Yb ytterbium 173.1	71	Lu lutetium 175.0
89	Ac actinium	90	Th thorium 232.0	91	Pa protactinium 231.0	92	U uranium 238.0	93	Np neptunium	94	Pu plutonium	95	Am americium	96	Cm curium	97	Bk berkelium	98	Cf californium	99	Es einsteinium	100	Fm fermium	101	Md mendelevium	102	No nobelium	103	Lr lawrencium

actinoids

